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Homework. Write Net Ionic Equations for All Precipitation Reactions. 1. Reactions Between Ions in Aqueous Solutions. 2. Precipitation Reaction Stoichiometry Practice. 1.

## **Advanced Chemistry - Unit 3-Reactions in Aqueous Solutions**

Reactions of metal ions in aqueous

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solution Chemistry A-level (7405) This resource (v1.4) represents colours of solutions and products (Specification reference 3.2.6 Reactions of ions in aqueous solution). Students are expected to describe: Metal Aqueous ion Action of NaOH Action of an excess of NaOH(aq) 3 Action of NH<sub>3</sub> (aq) Action of an excess

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## **A-level Chemistry Reactions of metal ions in aqueous solution**

Section 9.3 Reactions in Aqueous  
Solutions pages 299–308 Practice  
Problems pages 302–306 Write  
chemical, complete ionic, and net ionic  
equations for the following reactions  
that may produce precipitates. Use NR



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to indicate if no reaction occurs.

## **Chemical Reactions**

Summary Aqueous solutions can be classified as polar or nonpolar depending on how well they conduct electricity

Most chemical reactions are carried out in solutions, which are homogeneous

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mixtures of two or more substances. In a solution, a solute (the substance present in the lesser amount) is dispersed in a solvent (the substance present in the greater amount).

### **4.4: Types of Chemical Reactions in Aqueous Solution ...**

4.E: Reactions in Aqueous Solution

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(Exercises) These are homework exercises to accompany the Textmap created for "Chemistry: The Central Science" by Brown et al. 4.S: Reactions in Aqueous Solution (Summary) This is the summary Module for the chapter "Reactions in Aqueous Solution" in the Brown et al. General Chemistry Textmap.

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## **4: Reactions in Aqueous Solution - Chemistry LibreTexts**

Section 11.3 Reactions in Aqueous  
Solution 343 An ion that appears on both  
sides of an equation and is not directly  
involved in the reaction is called a  
spectator ion. When you rewrite an  
equation leaving out the spectator ions,

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you have the net ionic equation. The net ionic equation is an equation for a reaction in solution that shows

## **11.3 Reactions in Aqueous Solution**

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## **Chemistry 9.3 Reactions in Aqueous Solutions Flashcards ...**

replacement reactions. 9.3 Reactions in Aqueous Solutions MAIN Idea Double-replacement reactions occur between substances in aqueous solutions and produce precipitates, water, or gases. ChemFacts • Wood has to be heated to

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260°C ... 282 Chapter 9 • Chemical  
Reactions Section 99.1.1

## **Chapter 9: Chemical Reactions**

\_\_\_ 18. Will a precipitate form when 0.1 M aqueous solutions of  $\text{Ba}(\text{NO}_3)_2$  and  $\text{H}_2\text{CO}_3$  are mixed? If a precipitate does form, identify the precipitate and give the net ionic equation for the reaction. a.

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No precipitate forms. b.  $\text{BaCO}_3$   
2 precipitates.  $\text{Ba}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq})$   
 $\text{BaCO}_3(\text{s})$  c.  $\text{BaCO}_3$  precipitates.

## **AP Chem: Chapter 4 Practice Multiple Choice Questions**

Section 9.1. question. reactant. answer.  
A chemical substance that is present at  
the start of a chemical reaction.



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question. product. ... Reactions in  
Aqueous Solutions. answer. Section 9.3.  
question. aqueous solutions. answer.  
solutions in which water is the solvent.  
question. reactions that form  
precipitates.

**Chapter 9 Study Guide - McGraw Hill  
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Solid silver bromide and calcium nitrate are produced upon mixing an aqueous solution of 20.00 mL that is 0.0661 M silver nitrate with an aqueous solution of 35.00 mL that is  $\{eq\}4.724 \times 10^3$  ...

**Solid silver bromide and calcium nitrate are produced upon ...**

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Many reactions of this type involve the exchange of ions between ionic compounds in aqueous solution and are sometimes referred to as double displacement, double replacement, or metathesis reactions. These reactions are common in nature and are responsible for the formation of coral reefs in ocean waters and kidney stones

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in animals.

## **4.2 Classifying Chemical Reactions - Chemistry**

top. Test tube reactions of some transition metal aqueous ions. Iron(II) This exists in solution as the hexaaquairon(II) complex ion,  $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$ . It is a pale green solution, which

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slowly oxidises on standing in air to give a yellow/brown solution of hexaaquairon(III) ions.

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