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precipitate. bubbles. signs of a chemical reaction. subscript. Solid compound produced from 2 aqueous solutions

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during a chem.... Gas given off during a chemical reaction; one of the signs of.... change in heat/light, formation of a gas (bubbles), solid prec....

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The formulas of the reactants are on the
left, seperated by plus signs, and the
formulas of the products are on the
right, also ...

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another substance, often producing light or heat. reactant.

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oxidation-reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form oxidant + reductant →

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reduced oxidant + oxidized reductant.
The general forms of these five kinds of reactions are summarized in Table 11.6.1, along with examples of each.

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chemical reactions. release of a gas, color changes, formation of a precipitate, changes in heat and light. Energy. the ability to do work; different forms (heat, light electricity), 2 kinds: kinetic and potential. kinetic energy.

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Chapter 11: Chemical Reactions Study Guide. A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two. Small whole number ratios that are placed in front of the formulas in an equation in order to balance it.

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form two new compounds.

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substances, called reactants, are converted into one or more new substances, called products. The general form of the equation for such a process looks like this. (11.1.1) Reactants → Products

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To describe a chemical reaction. To calculate the quantities of compounds produced or consumed in a chemical reaction. As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano.

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Types of Chemical Reactions - 11.2

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is only one of the products of the reaction, and can thus weigh less than the original wood.

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